

Name: _____ Date: _____ Per.: _____

Ionic and Covalent Bonds Practice

For each of the following bonds:

- Determine if it will form an ionic or covalent bond
- If Ionic:
 - Draw a Lewis Dot structure for the valence shell of each element.
 - Draw an arrow (or more if needed) to show the transfer of electrons to the new element.
 - Write the charges on the ions.
 - Write the resulting chemical formula.
- If Covalent:
 - Draw the Lewis Dot structure for the compound.

• Use T-BO

a) Na + Cl

Ionic or Covalent:

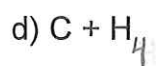
b) Cl + Cl

Ionic or Covalent:

c) Na + O

Ionic or Covalent:

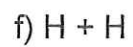
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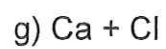
Ionic or Covalent:



Ionic or Covalent:

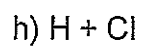


Ionic or Covalent:

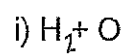


Ionic or Covalent:

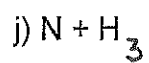
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Ionic or Covalent:



Ionic or Covalent:



Ionic or Covalent:



Ionic or Covalent:

Lewis Dots for Covalent Bonds

Draw Lewis dot structures for the following formulas.

1. H_2O	2. NH_2F	3. CHBr_3
4. Br_2	5. CH_2O	6. HCN
7. SiH_3F	8. CF_4	9. OBr_2
10. NH_2I	11. NH_3	12. SO_3
13. BF_3	14. HF	15. PCl_3